

01.08.2019

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De Mazenod College, Kandana

09 E I

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Final Term Test – Grade 13 - July 2019

Index Number

BIOLOGY I

Time : 3 hours

Answer all questions.

- Which is not a monosaccharide,
 - 1) Glucose
 - 2) Maltose
 - 3) RuBP
 - 4) Ribose
 - 5) Fructose
- Select the correct monomer of the given polymers,
 - 1) Inuline – Fructose
 - 2) Hemicellulose – Glucosamine
 - 3) Chitin – Pentose
 - 4) Pectin – Glucose
 - 5) Glycogen – Galacturonic acid
- A nucleotide that act as an oxidative agent,
 - 1) NADH
 - 2) ATP
 - 3) GTP
 - 4) NADP⁺
 - 5) ATP
- False statement about cell membrane,
 - 1) Cholesterol molecules provide flexibility
 - 2) It is selectively permeable
 - 3) It is about 7 mm thick
 - 4) It is compared to the fluid mosaic model.
 - 5) Some proteins have short branching carbohydrates.
- Correct statement about microscope.
 - 1) Visible light is passed back and forth through the specimen in light microscope.
 - 2) Electron beam is used in electron microscope.
 - 3) Resolution power of electron microscope is 0.2mm.
 - 4) Image is directly observed in electron microscopes.
 - 5) Dyes are used to stain the specimen in SEM.
- Correct statement about flagella.
 - 1) Made of microtubules.
 - 2) Basal bodies has 9 + 2 arrangement.
 - 3) Covered by cytoskeleton.
 - 4) Consists of microbodies.
 - 5) It is a short cylindrical structure.
- Which is not a significance of mitosis.
 - 1) Asexual reproduction.
 - 2) Growth and development.
 - 3) Maintains the genetic stability.
 - 4) Cell repair and replacement.



15. Incorrect statement about cells of ground tissue system of plant shoot.
- 1) Parenchyma cells are dead cells at maturity.
 - 2) These cells are specialized for storage and photosynthesis.
 - 3) Collenchyma cells are generally elongated
 - 4) Sclerite cell wall is thickened by lignin.
 - 5) Parenchyma cells possess large central vacuole.
16. Correct statement about secondary growth in plants
- 1) Occurs only in dicot plants.
 - 2) It increases the height of the plant.
 - 3) Xylem vessels achieve their conducting ability due to secondary growth.
 - 4) Periderm is the tissues that are interior to cork cambium.
 - 5) Periderm is impermeable to water and gases.
17. Select the correct deficiency symptoms in plants.
- 1) Fe - roots stunted and excessively branched.
 - 2) Cu - Light purple colour through our young leaves.
 - 3) Mg - Crinkled leaves.
 - 4) Ca - Crinkling of young leaves.
 - 5) P - Chlorosis between veins.
18. Select the correct example given for the respiratory structures
- 1) Body surface - Scorpions
 - 2) External gills - Aves
 - 3) Internal grills - Marien annelids
 - 4) Skin - Amphibia
 - 5) Book lungs - Some insects
19. Change that does not occur during ventilation of the lungs.
- 1) Pressure gradient is created between the atmosphere and lungs.
 - 2) During inhalation external intercoastal muscles contract.
 - 3) Exhalation is a passive process.
 - 4) During inhalation ribcage move up and out.
 - 5) Volume of lungs increase an visceral pleural relaxes.
20. Select the blood group of the donor when O(-) recipient is transfused.
- | | | |
|----------|---------|---------|
| 1) AB(-) | 3) O(-) | 5) A(-) |
| 2) AB(+) | 4) O(+) | |
21. Which is not a constituent in blood plasma
- 1) Hemoglobin
 - 2) Respiratory gases
 - 3) Albumin
 - 4) Antibodies
 - 5) Fibrinogen
22. False statement about human blood tissue.
- 1) Agglutinin is found on the surface of Red blood cells.
 - 2) Platelets are derived from epithelial cells of spleen.
 - 3) p^H of human blood is around 7.4
 - 4) Leukocytes develop immune responses in an infection
 - 5) Erythrocytes are biconcave disk - like cells.
23. A substance that does not secreted in to tubular filtrate in urine formation
- 1) H^+
 - 2) K^+
 - 3) Creatinine



5) Once spermatogenesis process is done, sperms are release into the fluid – filled cavity.

32. False statement about micro – organism

- 1) Mollicutes are prokaryotic organisms.
- 2) Viruses get into the host cell; they multiply and cause infections.
- 3) Unicellular protists are pleomorphic
- 4) Mutualistic fungi are found in lichens and mycorrhizae
- 5) *Nostoc* is free-living form of cyanobacteria that able fix atmospheric N_2

33. Correct statement about food spoilage

- 1) Accumilation of toxins is a chemical change
- 2) Rancidity is caused by proteolytic bacteria
- 3) Food intoxication is caused by *Shigella*.
- 4) Aflatoxins are produced by *Aspergillus flavus*
- 5) Relative humidity does not affect on food spoilage.

34. Select the correct order in translation process.

- a) Amino acids are added to the C – terminus of the growing polypeptide chain.
- b) Release of the t RNA in the P site
- c) t RNA carries methionine
- d) A site is then ligned with the next codon.
- e) Small subunit of ribosome binds to m RNA

- 1) a b d c e
- 2) e a c b d
- 3) e c d a b
- 4) a b c d e
- 5) e c a b d

35. Which is not an application of DNA sequencing.

- 1) To reveal evolutionary relationships within and among species.
- 2) Diagnosis of cancer
- 3) To determine the person on carrier or not regarding family inherit disease.
- 4) To understand the functions of DNA
- 5) To make recombinant vectors

36. False statement about human upper limb

- 1) Carpel bones are arranged in two rows
- 2) Radius articulate with three carpel bones at its distal end to form wrist joint
- 3) Opposable thumb permits precision grip
- 4) Ulna articulate with the distal end of humerus at elbow joint
- 5) Humerus forms a complete ball socket joint in glenoid cavity

37. Which is not a function of human skeletal system

- 1) Production of blood cells
- 2) Support
- 3) Movement
- 4) Protection
- 5) Transportation

38. Select the false statement about human genetic disorders

- 1) In males x – linked recessive disorders are expressed only at their homozygous genotype
- 2) Red – green colour blindness is caused by x – linked recessive disorder
- 3) Haemophilia is caused by x – linked recessive disorder
- 4) Cystic fibrosis is an autosomal recessive disorder
- 5) Sickle cell disease is caused by single gene mutation.

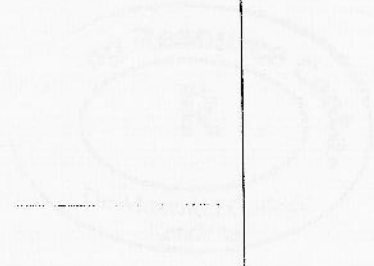
9. Effects of global warming

- A. Degradation of coral reefs
- B. Sea level rising
- C. Depletion of ozone layer
- D. Extreme water events
- E. Photochemical smog

10. Which is / are contributory factors for desertification

- A. Acid rains
- B. Over – exploitation of soil
- C. Uncontrolled mining
- D. Excessive use of agro chemicals
- i. Emission of green house gases.

BIOLOGY



DE MAZENOD COLLEGE – KANDANA
THIRD TERM TEST – 2019 JULY
BIOLOGY – GRADE 13

01.08.2019

3 HOURS

Part A – Structured Essay

Answer the questions using the given space.

1. A)

i. Name the elements that found as 96% in living matter.

.....

ii. Why water is important fo life?

.....

.....

iii. Name a Nitrogen containing polysaccharide.

.....

iv. Name a Structural polysaccharide found in prokaryotes.

.....

vi. (a) What is meant by resolution power of microscope?

.....

(b) What is the resolution of light microscope?

.....

vii. (a) Draw and label the ultra-structure of mitochondria.



(b) What are the functions of the above organelle?

.....
.....

B)

i. What is meant by respiratory quotient?

.....

ii. What are the major differences of lactic acid fermentation and Alcoholic fermentation?

.....
.....
.....

iii. What are the major processes that take place in CO₂ fixation in C₃ plants?

.....
.....
.....

iv. What are the factors that affect on the rate of photosynthesis?

.....
.....
.....
.....

v. What are allosteric regulators?

.....

vi. What are the changes occur in 'S-' phase of interphase in mitotic cell division.

.....
.....

(C)

i. Name the significances of mitosis

.....
.....
.....
.....

ii. Briefly explain what is meant by crossing over.

.....
.....
.....

2. A)

i. What is meant by radial transport?

.....

ii. Name the different types of routes used in the radial transport.

.....
.....
.....

iii. What is meant by nutrition in plants?

.....
.....

iv. What are the needs of nutrition in plants?

.....
.....
.....

v. Why essential elements are needed for plants?

.....
.....

vi. Name deficiency symptoms of Ca in plants.

.....
.....

vii. Name the source and the form of absorption of P into plants.

Source: -

Form: -

viii. Name the soil microbial interaction in effective uptake of P into plants.

.....

B)

i. Name the key steps in the Nitrogen cycle.

.....
.....
.....
.....

ii. what are the applications of micro-organisms in agriculture?

.....
.....
.....

iii. Name the indicator organism used to test the quality of drinking water.

.....

iv. State the characteristic features of the above mentioned organism.

.....
.....
.....

v. Name two water born diseases

.....
.....



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SI - abar
Biology

vi. What are the adverse effects of discharging large amount of waste water into natural water bodies?

.....
.....
.....

(C) i. What is the functional unit of liver?

.....

ii. What are the functions of liver related to digestion?

.....
.....
.....

iii. (a) Name the hormones released by the duodenum.

.....

(b) What are the functions of the above hormones.

.....
.....
.....

3. A) i. (a) Name the efferent components of peripheral nervous system

a-.....

b-.....

b) Write the function of each of the above-mentioned components.

a-.....

b-.....

ii. Name the functions of human cerebellum

.....
.....
.....

iii. What are the factors required to maintain the resting membrane potential?

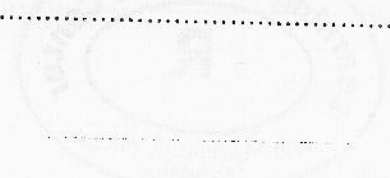
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.....

iv. What are the phases of action potential?

.....
.....
.....

v. Name the different types of photoreceptors in human eye and state the visual pigment in those receptors

.....
.....



i.(a) Name the tissues that form the dermis of the human skin.

.....

(b) Name the main types of cells found in the above-mentioned tissue.

.....

(c) Name 4 structures found in the dermis.

.....
.....

(d) State functions of human skin.

.....
.....
.....
.....

ii. Name a non-trophic hormone secreted by anterior pituitary.

.....

iii. What is the function of above-mentioned hormone?

.....

i. What is the function of thymosin?

.....

ii. What are antibodies?

.....

iii. Name basic features of acquired immunity.

.....
.....
.....

iv. Name an important inflammatory signalling molecule.

.....

v. What are the signs and symptoms of inflammation?

.....
.....

4. A) i. What is meant by dihybrid test cross?

.....
.....

ii. 1280 plants were obtained in F1 generation in a cross between genotypes Aa BB cc X Aa bb CC.

(a) Find the probability of obtaining the AaBbcc in their F1 generation

.....
.....
.....

(b) Find the number of organisms that possess the Genotype AaBbcc?

.....

iii. Red-green colour blindness is caused by sex linked recessive allele. In a family of normal parent's son is colour blind and their daughter is normal for colour blindness. write the genotypes of family members?

(Dominant allele = C recessive allele = c)

Mother

Son

Father

Daughter

iv. What is meant by mutation?

.....
.....

v. What are the different types of gene mutations?

.....
.....
.....

B) i. Name three genetic disorders that occur due to chromosomal mutations.

.....
.....
.....

ii. State major steps and techniques used to make a recombinant vector?

.....
.....
.....
.....
.....
.....

iii. What is meant by cloning site?

.....
.....

iv. Name an international agreement established in order to safeguard consumers, society and environment due to potential risks of GMO s?

.....

v. What is the objective of the above mentioned agreement?

.....
.....

vi. What are the functions of DNA polymerase in DNA replication?

.....
.....
.....

vii. What are the three main steps in elongation of translation?

(C) i. State 5 characteristic features of the vegetation in desert biome in tropics?

- 1
- 2
- 3
- 4
- 5

ii. Name the values of Biodiversity?

- 1
- 2
- 3
- 4
- 5

iii. Write two human activities leads to desertification?

- 1
- 2



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De Mazenod College Kandana

01 E II

අවසන් වාර පරීක්ෂණය - 13 ශ්‍රේණිය - 2019 ජූලි
Final Term Test – Grade 13 – July 2019

Index Number

BIOLOGY I I

Time : 3 hours

❖ Part B – Essay

Answer four questions only.

PART B – ESSAY

1. Explain the process of osmoregulation in human body.
2. Briefly explain the lung ventilation mechanism and its regulation
3.
 - i. Explain the vitality of water in living matter with respect to its major properties.
 - ii. Explain the stomatal transpiration in higher plants.
4.
 - i. Briefly explain the ultrastructure of chloroplast in eukaryotic cell.
 - ii. Explain the C_3 cyclic pathway of CO_2 fixation in photosynthesis.
5.
 - i. Explain sterilization techniques used in laboratory preparation of microbial culture
 - ii Briefly explain the physical and chemical changes that take place in microbial food spoilage.
6. Write short notes.
 - i. Cardiac conduction system.
 - ii. Acid rains
 - iii. Gross structure of human ovary

Part B - Essay

Answer All Questions

5. (a) (i) Deduce the relationship between the relative molecular mass (M) and the density (d) of a particular gas by using the ideal gas equation.
- (ii) A mixture of gases contains 75% of N_2 gas and 25% of O_2 gas by volume. The pressure of the mixture of gases is 1.00×10^5 Pa and the temperature is 300 K. Assuming ideal behavior, calculate the following.
- The partial pressure of O_2 in this mixture of gases
 - The relative molecular masses relevant for this mixture of gases. (the relative atomic masses of N and O are 14.0 and 16.0 respectively)
 - The density of the mixture of gases.

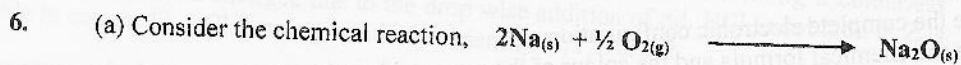
- (iii) Various mixtures consisting with finely powdered Al, Mg and Zn were prepared, each weighing 100 g. Masses of different metals used for the preparation were given below.

Mixture	Mass of the metals / (g)			Total mass/ (g)
	Al	Zn	Mg	
P	21.6	13.0	65.4	100.0
Q	27.0	52.0	21.0	100.0
R	32.4	65.0	2.6	100.0

(Al = 27, Mg = 24, Zn = 65)

However the labeling of the samples was found to be inaccurate. Using balanced chemical equations and proper calculation, how would you identify the samples separately?

- (iv) Speeds of six gaseous molecules of an ideal gas measured at -250°C were given as, 2.0 m s^{-1} , 2.2 m s^{-1} , 2.6 m s^{-1} , 2.7 m s^{-1} , 3.3 m s^{-1} and 3.5 m s^{-1} respectively. Using these values calculate the following.
- Average speed of a gas molecule
 - Root mean square speed
 - Molar mass of the gaseous species
 - Density of the gas at 10 Pa pressure



Some of the thermo chemical data given relevant to the above reaction are given below. (at 25°C .)

Atomization enthalpy of $\text{Na}_{(s)}$	= 107 kJ mol^{-1}
First ionization enthalpy of $\text{Na}_{(g)}$	= 494 kJ mol^{-1}
Atomization enthalpy of $\text{O}_{2(g)}$	= 249 kJ mol^{-1}
First electron affinity of $\text{O}_{(g)}$	= -149 kJ mol^{-1}
Second electron affinity of $\text{O}_{(g)}$	= 798 kJ mol^{-1}
Lattice enthalpy of $\text{Na}_2\text{O}_{(s)}$	= 2522 kJ mol^{-1}



- I. Write chemical equations to express above enthalpy changes.
 - II. Calculate the enthalpy change of formation of $\text{Na}_2\text{O}_{(s)}$ by constructing an appropriate enthalpy level diagram.
- b)
- Standard enthalpy change of hydration of $\text{Li}^+_{(g)}$, $\text{Na}^+_{(g)}$ and $\text{Cl}^-_{(g)}$ are -499 kJ mol^{-1} , -390 kJ mol^{-1} and -381 kJ mol^{-1} respectively. And standard enthalpy change of lattice of $\text{LiCl}_{(s)}$ and $\text{NaCl}_{(s)}$ are -864 kJ mol^{-1} and -780 kJ mol^{-1} .
- I. Calculate the standard enthalpy change of dissolution of $\text{LiCl}_{(s)}$ and $\text{NaCl}_{(s)}$
 - II. State, whether the temperature of the system increase or decrease, when $\text{LiCl}_{(s)}$ is dissolving in water.
- c) Calculate the enthalpy of formation of $\text{NaCl}_{(s)}$ from the following data

Reaction	$\Delta H^{\circ}_{298} / \text{kJ mol}^{-1}$
$\text{NaOH}_{(aq)} + \text{HCl}_{(aq)} \rightarrow \text{NaCl}_{(aq)} + \text{H}_2\text{O}_{(l)}$	-57.3
$\text{HCl}_{(g)} + \text{aq} \rightarrow \text{HCl}_{(aq)}$	-71.9
$\text{H}_{2(g)} + \frac{1}{2} \text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$	-285.9
$\frac{1}{2} \text{H}_{2(g)} + \frac{1}{2} \text{Cl}_{2(g)} \rightarrow \text{HCl}_{(g)}$	92.3
$\text{Na}_{(s)} + \frac{1}{2} \text{O}_{2(g)} + \frac{1}{2} \text{H}_{2(g)} \rightarrow \text{NaOH}_{(aq)}$	-425.6
$\text{NaCl}_{(s)} + \text{water} \rightarrow \text{NaCl}_{(aq)}$	+3.9

7. (a) "L" is an element that belongs to the p block of the periodic table. In the group that L belongs, there are metals, non metals as well as metalloids. The first member of the group is distinctive as it has a tendency to form a large amount of compounds. L is the second member of this group.
- i. Identify L
 - ii. Write the complete electronic configuration of L
 - iii. Write the chemical formula and the colour of the most stable oxide of L.
 - iv. L is very important in the manufacture of integrated circuits and transistors. State another important use of L
- (b) Consider the oxidation of SnC_2O_4 by KMnO_4 in the dil. H_2SO_4 medium.
- i. Write the reduction half reaction
 - ii. Write the oxidation half reaction
 - iii. Give the balanced redox reaction

(c) You are supplied with a mixture consisting of barium carbonate, dolomite, potassium carbonate and silicon dioxide. Explain how you would attempt to determine quantitatively each of the constituents present in this mixture.

(d) When 20.0 g of a white solid X is heated, 4.4 g of an acidic gas A and 1.8 g of a neutral gas B are evolved, leaving behind a solid residue of mass 13.8 g. A turns lime water milky and B condenses into a liquid which changed anhydrous cupric sulphate blue. The aqueous solution Y is alkaline to litmus and gives 19.7g of white precipitate Z with barium chloride solution. Z gives carbon dioxide with an acid. Identify A, B, X, Y and Z. (C= 12, H= 1, O = 16, Ba= 137)

8. (a) State all possible reactions (balanced) which can be occurred when a piece of Na is exposed to air.

(b) How would you distinguish between the below mentioned pairs?

- i. Li_2CO_3 and Na_2CO_3
- ii. Na_2CO_3 and MgCO_3
- iii. LiNO_3 and NaNO_3
- iv. KNO_3 and $\text{Ba}(\text{NO}_3)_2$
- v. Na_2CO_3 and NaHCO_3
- vi. LiOH and NaOH

(c) What will happen to the water solubility of below mentioned species formed by alkaline earth metals down the group?

- i. Carbonates
- ii. Sulphates
- iii. Oxalates
- iv. Hydroxides
- v. Chromates

(d) Assume that you have provided a mixture containing KNO_3 and K_2CrO_4 . You are provided $\text{Ba}(\text{NO}_3)_2$ and CH_3COOH as the only chemical reagents. How would you confirm the presence of KNO_3 using the given reagents. (Heating is allowed)

(e) X is an element of the main group of the periodic table. Nitrate of the element X is soluble in water forming a solution Y. White precipitate Z can be obtained due to the drop wise addition of dil. NaOH to solution Y. Z is soluble in excess NaOH forming a colourless solution P. Z can be re-emerged due to the drop wise addition of dil. HCl to solution P. Z is soluble in excess dil. HCl forming solution Y. Identify X, Y, Z and P.

(f) A mixture of zinc and aluminum weighing 1.67 g was completely dissolved in dilute sulphuric acid and evolved 1.69 L of hydrogen gas measured at 273 K and 1 atm pressure. What was the weight of aluminum in the original mixture? (Assume the ideal behavior of hydrogen gas, Al= 27, Zn = 65)

PERIODIC TABLE OF THE ELEMENTS

1 H Hydrogen 1.00794																	18 He Helium 4.002602		
3 Li Lithium 6.941	4 Be Beryllium 9.012182																	9 F Fluorine 18.9984032	10 Ne Neon 20.1797
11 Na Sodium 22.98976928	12 Mg Magnesium 24.304																	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.0983	20 Ca Calcium 40.078	21 Sc Scandium 44.955912	22 Ti Titanium 47.88	23 V Vanadium 50.9415	24 Cr Chromium 51.9961	25 Mn Manganese 54.938044	26 Fe Iron 55.845	27 Co Cobalt 58.933195	28 Ni Nickel 58.6934	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.64	33 As Arsenic 74.9216	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.798		
37 Rb Rubidium 85.4678	38 Sr Strontium 87.62	39 Y Yttrium 88.90584	40 Zr Zirconium 91.224	41 Nb Niobium 92.90638	42 Mo Molybdenum 95.94	43 Tc Technetium 98	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.9055	46 Pd Palladium 106.3635	47 Ag Silver 107.8682	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.757	52 Te Tellurium 127.6	53 I Iodine 126.90548	54 Xe Xenon 131.29		
55 Cs Cesium 132.90545196	56 Ba Barium 137.327	57-71 Lanthanoids	72 Hf Hafnium 178.49	73 Ta Tantalum 180.94788	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.222	78 Pt Platinum 195.084	79 Au Gold 196.966569	80 Hg Mercury 200.59	81 Tl Thallium 204.3833	82 Pb Lead 207.2	83 Bi Bismuth 208.9804	84 Po Polonium 209	85 At Astatine 210	86 Rn Radon 222		
87 Fr Francium 223	88 Ra Radium 226	89-103 Actinoids	104 Rf Rutherfordium 261	105 Db Dubnium 262	106 Sg Seaborgium 263	107 Bh Bohrium 264	108 Hs Hassium 265	109 Mt Meitnerium 266	110 Ds Darmstadtium 267	111 Rg Roentgenium 268	112 Cn Copernicium 269	113 Nh Nihonium 270	114 Fl Flerovium 271	115 Lv Livermorium 272	116 Ts Tennessine 273	117 Og Oganesson 274	118 Uu Ununseptium 275		
57 La Lanthanum 138.90547	58 Ce Cerium 140.12	59 Pr Praseodymium 140.90766	60 Nd Neodymium 144.242	61 Pm Promethium 144.9126	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.92534	66 Dy Dysprosium 162.50014	67 Ho Holmium 164.930329	68 Er Erbium 167.259	69 Tm Thulium 168.93032	70 Yb Ytterbium 173.05448	71 Lu Lutetium 174.96706					
89 Ac Actinium 227	90 Th Thorium 232.0377	91 Pa Protactinium 231.03688	92 U Uranium 238.02891	93 Np Neptunium 237.04817	94 Pu Plutonium 244.06422	95 Am Americium 243.06138	96 Cm Curium 247.07537	97 Bk Berkelium 247.0703	98 Cf Californium 251.07958	99 Es Einsteinium 252.08322	100 Fm Fermium 257.1035	101 Md Mendelevium 258.1051	102 No Nobelium 259.10636	103 Lr Lawrencium 260.1073					